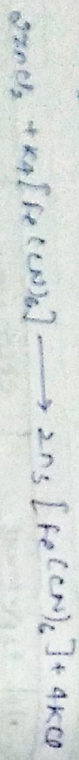


Reactions 1



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EXPERIMENT - 19.

Aim 1 to analyze the cation and anion in the given salt.

Observations 1

Sr.No	Experiment	Observations	Inference
1.	Salt + dil. H ₂ SO ₄	No rxn.	CO ₃ ²⁻ , CH ₃ COO ⁻ absent
2.	Salt + conc. H ₂ SO ₄	Pungent smell	Cl ⁻ possible
3.	Bring a glass rod dipped in NH ₄ OH on the mouth of the test tube.	White fumes come out	Cl ⁻ confirm
4.	Salt + NaOH + Heat	No smell	ZnO ppt absent
5.	0.5% dil. HCl	No reaction	I ⁻ ppt absent
6.	Pass H ₂ S gas in the soln.	No ppt.	I ⁻ ppt absent
7.	Pass H ₂ S gas and NH ₄ Cl + NH ₄ OH	No ppt.	S²⁻ ppt absent
8.	0.5% + NH ₄ OH. (Excess)		Zn ²⁺ confirm

Result 1 Given salt contains acidic sulfate Cl⁻ and basic soluble Zn²⁺

Signature

Teacher's Signature

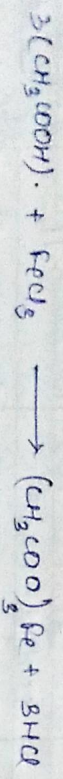
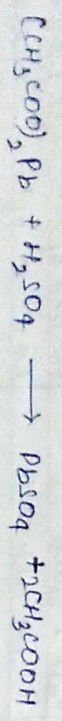
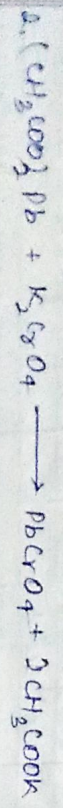
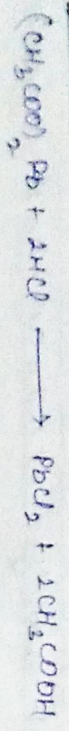
EXPERIMENT-20.

AIM - to analysis one cation and anion in the given salt.

Sr No.	Experiment	Observation	Inferences
1.	Salt + 2 or 3 drops of conc. H_2SO_4	Pungent smell.	Cl^- possible
2.	Bring a glass rod dipped in NH_4OH	white fumes	Cl^- confirm
3.	salt + $NaOH$	No smell of ammonia	NH_4^+ absent
4.	O.S. + dil. HCl	No white ppt.	I^{st} gp absent
5.	O.S. + dil. HCl + boil + NH_4Cl + NH_4OH	No brown ppt.	II^{rd} & III^{rd} gp. absent
6.	O.S. + NH_4Cl + NH_4OH + boil + H_2S	No white ppt.	IV^{th} gp absent
7.	O.S. + NH_4OH + CH_3COOH	No white ppt.	V^{th} gp absent
8.	O.S. + NH_4OH + Na_2HPO_4	white ppt.	Mg^{2+} confirm

Result - The given salt contain Mg^{2+} cation and Cl^- anion.

Reactions +



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EXPERIMENT - 21

Aim - To analyze a cation and an anion in the given salt,

Observations

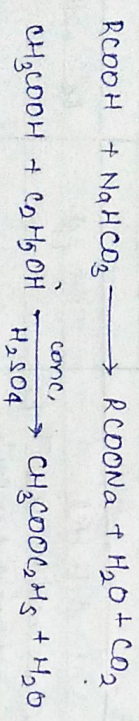
1. Salt + dil. HCl	white ppt.	Pb ²⁺ possible
2. Acetic acid in the sol ⁿ of,		
3. divide into two parts		
1) Add KI	yellow ppt.	Pb ²⁺ confirm
2) Add K ₂ CrO ₄	yellow ppt	Pb ²⁺ confirm
4. Salt + dil. H ₂ SO ₄ + heat	small of vinegar	CH ₃ COO possible
5. Salt + concen. + neutral FeCl ₃	red colour (the blood)	CH ₃ COO confirm

Result - The given salt contain CH₃COO⁻ anion and Pb²⁺ cation confirm.

Answer

Teacher's Signature : _____

Reactions :-



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EXPERIMENT - 2A

Objective :- Identify the functional gp present in the given O.C.

Physical state :- liquid

colour :- colourless

smell :- Vinegary

solubility in water :- water soluble.

Flame test :- It burns with a non-sooty flame therefore it is a aliphatic compound.

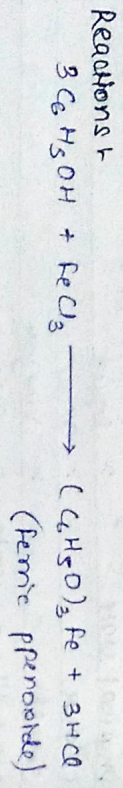
Experiment	Observation	Inference
Organic compound + H ₂ O + NaHCO ₃	Bubbly effervescence	-COOH gp present

Esterification :- when 1 mole of acid and 1 mole of alcohol is heated with 2 drops of conc. H₂SO₄ a fragrant smell comes out due to formation of ester.

Result :- Given O.C. contain -COOH gp.

Garuru

Teacher's Signature : _____



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EXPERIMENT - 23

Objective 1 To identify functional gp. present in given organic compound.

Colour Reddish brown

Odour Characteristic odour

Water solubility Insoluble in water

Physical state Liquid

Aromaticity test It burns with sooty flame so, it is aromatic.

Functional gp test

Experiment	Observation	Inference
D.C. + FeCl ₃ + excess + H ₂ O	Purple colouration	Phenolic - OH gp. present

Result - The given organic compound contains phenolic - OH gp.

Teacher's Signature

Teacher's Signature : _____